

RESERVE COPY

PATENT SPECIFICATION

1,152,5

DRAWINGS ATTACHED.

1,152,575



Date of Application and filing Complete Specification:
16 May, 1966. No. 21654/66.

Application made in United States of America (No. 456,219) on
17 May, 1965.

Complete Specification Published: 21 May, 1969.

© Crown Copyright 1969.

Index at Acceptance:—C2 C(3A10B2C, 3A10B5C, 3A10B5G1, 3A10D2, 3A10D5C, 3A10D5G1)
Int. Cl.:—C 07 c 63/00.

COMPLETE SPECIFICATION.

Purification of Aromatic Polycarboxylic Acids.

PATENTS RULES, 1968

SPECIFICATION NO. 1,152,575

The following amendment was allowed under Rule 94(3) on 2nd July 1969:—

Page 1, in the heading, after "Application made in United States of America (No. 456,219 May, 1965)" insert "Patent of Addition to Patent No. 994,769 dated 25 March, 1964."

THE PATENT OFFICE,
11th August 1969

D 1

diols to produce super polyesters. Though the process of the invention is mainly described herein as applied to the purification of crude terephthalic acid to produce terephthalic acid of fiber-grade quality, it is not limited thereto and it includes the purification of other polycarboxylic aromatic acids such as, for example, trimesic, isophthalic, naphthalene dicarboxylic, trimellitic and mellitic acids.

High molecular weight polyesters of terephthalic acid with various diols find extensive use as synthetic fibers and film. These super polyesters, first described in British Patent Specification No. 578,709, have heretofore been prepared from dimethylterephthalate, which is trans-esterified with the appropriate diol, such as ethylene glycol, and then polycondensed to form the super polyester. Polyester preparation via dimethylterephthalate has been considered an essential step by reason of the exceptionally high purity requirements imposed on the polyester.

With the advent of improved processes for the manufacture of terephthalic acid, much attention has been directed to obtain-

tained from the oxidation of paraxylene or other para-disubstituted alkyl benzene, is known to be deleterious with respect to polyester quality. Second, unidentified color bodies, possibly of the benzil or fluorenone structure, are usually present as trace by-products of most terephthalic acid production processes, and yield off-color polyesters. Any method of purifying terephthalic acid must reduce or eliminate both the 4-carboxybenzaldehyde (4-CBA) and the other impurities such as color bodies.

In accordance with the invention, polycarboxylic acid of a purity suitable for direct esterification with diols to produce films and fibers is obtained from impure polycarboxylic aromatic acid containing aldehyde and other impurities by catalytically hydrogenating the aldehyde and other impurities in the impure acid by contacting a solution of the impure acid in water in the liquid phase at an elevated temperature and pressure and in the presence of hydrogen with a platinum group metal-containing catalyst under mild hydrogenating conditions (as hereinafter defined) and recovering from the thus contacted solution purified

PATENT SPECIFICATION

1,152,575

DRAWINGS ATTACHED.

1,152,575



Date of Application and filing Complete Specification:
16 May, 1966. No. 21654/66.

Application made in United States of America (No. 456,219) on
17 May, 1965.

Complete Specification Published: 21 May, 1969.

© Crown Copyright 1969.

Index at Acceptance:—C2 C(3A10B2C, 3A10B5C, 3A10B5G1, 3A10D2, 3A10D5C, 3A10D5G1).

Int. Cl.:—C 07 c 63/00.

COMPLETE SPECIFICATION.

Purification of Aromatic Polycarboxylic Acids.

We, STANDARD OIL COMPANY, a Corporation organised and existing under the laws of the State of Indiana, United States of America, of 910 South Michigan Avenue, Chicago, 60680, Illinois, United States of America, do hereby declare the invention, for which we pray that a patent may be granted to us, and the method by which it is to be performed, to be particularly described in and by the following statement:—

This invention relates to the purification of polycarboxylic aromatic acids, and more particularly concerns the preparation of polycarboxylic aromatic acid having a purity sufficient for direct esterification with diols to produce super polyesters. Though the process of the invention is mainly described herein as applied to the purification of crude terephthalic acid to produce terephthalic acid of fiber-grade quality, it is not limited thereto and it includes the purification of other polycarboxylic aromatic acids such as, for example, trimesic, isophthalic, naphthalene dicarboxylic, trimellitic and mellitic acids.

High molecular weight polyesters of terephthalic acid with various diols find extensive use as synthetic fibers and film. These super polyesters, first described in British Patent Specification No. 578,709, have heretofore been prepared from dimethylterephthalate, which is trans-esterified with the appropriate diol, such as ethylene glycol, and then polycondensed to form the super polyester. Polyester preparation via dimethylterephthalate has been considered an essential step by reason of the exceptionally high purity requirements imposed on the polyester.

With the advent of improved processes for the manufacture of terephthalic acid, much attention has been directed to obtain-

ing polyesters by direct esterification of terephthalic acid with the diol. This has manifest advantages of simplicity and economy as compared with the dimethylterephthalate route. As yet, however, there remains serious difficulty in obtaining terephthalic acid of suitable purity. Unless the initial terephthalic acid is virtually completely free from extraneous contaminants, the polyester will have too low a melting point and will be of unsatisfactory color.

It is believed that terephthalic acid impurities are of two types. First, the compound 4-carboxybenzaldehyde, an intermediate formed when terephthalic acid is obtained from the oxidation of paraxylene or other para-disubstituted alkyl benzene, is known to be deleterious with respect to polyester quality. Second, unidentified color bodies, possibly of the benzil or fluorenone structure, are usually present as trace by-products of most terephthalic acid production processes, and yield off-color polyesters. Any method of purifying terephthalic acid must reduce or eliminate both the 4-carboxybenzaldehyde (4-CBA) and the other impurities such as color bodies.

In accordance with the invention, polycarboxylic acid of a purity suitable for direct esterification with diols to produce films and fibers is obtained from impure polycarboxylic aromatic acid containing aldehyde and other impurities by catalytically hydrogenating the aldehyde and other impurities in the impure acid by contacting a solution of the impure acid in water in the liquid phase at an elevated temperature and pressure and in the presence of hydrogen with a platinum group metal-containing catalyst under mild hydrogenating conditions (as hereinafter defined) and recovering from the thus contacted solution purified

acid containing substantially less aldehyde and other impurities by crystallization employing controlled-rate evaporative cooling. By this process, there is a chemical reduction of the aldehyde as well as a significant destruction of the color body impurity. Surprisingly, when the evaporation rate of water solvent from said solution is controlled, as by maintaining back-pressure, to a rate below the spontaneous uncontrolled vapor release rate, the purity of the crystallized acid is much increased.

Thus the invention provides a process as set forth in the preceding paragraph for purifying aromatic polycarboxylic acid produced by liquid phase catalytic oxidation of polyalkyl aromatic hydrocarbons to remove undesirable aldehyde and other impurities therefrom. The process is particularly well suited for purification of aromatic dicarboxylic acids as, for example, in the production of fiber-grade terephthalic acid from crude terephthalic acid. The platinum group metal-containing catalyst suitably consists essentially of charcoal having supported thereon 0.01 to 1.0% by weight of platinum group metal. Preferably the charcoal support is one having a surface area in the range of 1,000 to 2,000 square meters per gram. Particularly desirable metals for use as the catalyst are platinum and palladium, preferably palladium.

The process of the present invention may be applied to the purification of crude aromatic polycarboxylic acid obtained by catalytic liquid phase oxidation of a polyalkyl aromatic hydrocarbon with molecular oxygen in the presence of a heavy metal oxidation catalyst, said aromatic hydrocarbon having at least two nuclear alkyl hydrocarbon substituents whose carbon attached to the nuclear aromatic carbon has at least one hydrogen atom, which crude aromatic polycarboxylic acid has an aromatic polycarboxylic acid content of at least 99.0% by weight, preferably 99.5 wt. %, and which has as its principal impurity a carboxy aromatic aldehyde corresponding to said aromatic polycarboxylic acid, to obtain a purified acid containing less than 125 ppm of said aldehyde. The catalytic hydrogenation step may suitably be carried out at a temperature of from 450 to 600°F.

According to a preferred embodiment of the present invention, there is provided a method of producing fiber-grade terephthalic acid from crude terephthalic acid obtained by catalytic liquid phase oxidation of paraxylene with molecular oxygen in the presence of a heavy metal oxidation catalyst, which crude terephthalic acid has a terephthalic acid content of at least 99.0% by weight, preferably 99.5 wt. %, and which has as its principal impurity 4-carboxybenzaldehyde, which method comprises contact-

ing a solution of said crude terephthalic acid in water in the liquid phase at a temperature in the range of 450 to 600°F. in the presence of hydrogen with a platinum group metal-charcoal hydrogenation catalyst for a time sufficient to effect substantial reduction of said aldehyde to para-toluic acid and recovering fiber-grade terephthalic acid by crystallization from the contacted solution by controlled-rate evaporation of water from the solution.

When terephthalic acid, or other aromatic dicarboxylic acid to be employed in super polyester production, is purified by the process of the present invention, it is said to be of "fiber-grade" quality. The term "fiber-grade" does not denote a quantitative degree of purity, but rather describes a terephthalic acid which is sufficiently free from 4-carboxybenzaldehyde and other impurities so as to yield a super polyester upon direct esterification with a diol which is satisfactory for the intended purpose. Thus, terephthalic acid polyester intended for transparent film or white fiber will require significantly higher purity terephthalic acid than polyesters intended for tire reinforcing cord. The suitable 4-carboxybenzaldehyde (4-CBA) content of terephthalic acid for use in producing colorless film and fiber is less than 125 parts per million (ppm) while for tire cord it may be as high as 500 parts per million or even higher. The p-toluic acid and 4-CBA contents of the fiber-grade terephthalic acid are each suitably less than 125 ppm, desirably less than 75 ppm, preferably less than 25 ppm, ideally essentially zero. The content of 4-CBA is more critical than p-toluic acid.

The process of the present invention has particular application to purification of terephthalic acid produced by the liquid phase air (molecular oxygen) oxidation of paraxylene using a heavy metal and bromine as catalyst as described in British Patent Specification No. 807,091. The process of the present invention may also be used to advantage for purification of terephthalic acid from other processes for the catalytic liquid phase oxidation of para-dialkylbenzenes with molecular oxygen in the presence of heavy metal oxidation catalyst, such as those promoted with acetaldehyde or methyl ethyl ketone, for the terephthalic acids produced by these oxidation processes also contain 4-CBA impurity. Terephthalic acid from any source which contains 4-carboxybenzaldehyde and which is yellowish in color, can be converted to fiber grade terephthalic acid by the process of this invention.

The process of the invention is conducted at elevated temperature and pressure. By reason of its low solubility in water, terephthalic acid requires either large vol-

umes of water or high temperatures in order for the desired terephthalic acid production quantity to be put into solution. For reasons of economic equipment design and process operation, it is therefore desirable to conduct the process within the range of 392 to 700°F., although lower or higher temperatures may be used in particular circumstances. The most advantageous temperature range is about 450 to 600°F., e.g. 464 to 550°F. The quantity of water needed to dissolve the terephthalic acid at various temperatures may be estimated from the table below:

	Terephthalic Acid, g./100 g. H ₂ O	Temperature, °F for solution
15	1	365
	5	401
	10	468
20	20	498
	30	522

Ordinary process water is not completely satisfactory for use as the aqueous solvent in the process of the present invention where palladium is used as the catalyst, nor will ordinary distilled water do. It has been discovered that the palladium catalyst is poisoned by dissolved minerals, such as for example copper and other metals, therefore it is critical to the useful life of the platinum group metal catalyst used to effect the particular hydrogenation of this invention that such minerals be removed from the water to be employed as solvent. Advantageously the water, conveniently steam condensate, is demineralized to a total dissolved solids content of less than 10 parts per million parts of water (ppm), preferably to less than 5 ppm, optimally to less than 1 ppm. The water may be demineralized by any suitable process known to the art, such as for example by passing the water through a bed of chelating resin. The useful life of the palladium catalyst is greatly increased when minerals, and in particular copper, are removed. It is surprising that copper poisons the palladium since copper in various forms is known to be a hydrogenation catalyst.

Pressure conditions for the process of this invention depend upon the temperature at which this process is conducted. Since the temperatures at which significant amounts of the impure terephthalic acid may be dissolved in water are substantially above the normal boiling point of water, and since the hydrogenation section of the process of this invention is to be carried out with the solvent in the liquid phase, the pressure will necessarily be substantially above atmospheric pressure.

It is preferred to trickle the liquid solution of acid through a bed of the catalyst

because lower hydrogen partial pressure or hydrogen driving force is required than is required when the catalyst bed is operated liquid full. Either a static hydrogen atmosphere or a flow, concurrent or countercurrent, of hydrogen through the catalyst chamber may be maintained. Lower hydrogen partial pressures are required for the trickle or percolation method of conducting the hydrogenation section of the process of this invention because there is provided a thin film of the aqueous solution of the impure terephthalic acid on the catalyst particles and thus a lower hydrogen driving force is needed for the hydrogen to dissolve and diffuse through the thin liquid layer and reach the catalyst. For such percolation method of conducting the hydrogenation a continuous flow or atmosphere of hydrogen is not essential. However, for maximum hydrogenation rates it is beneficial to dissolve at least some hydrogen into the solution, conveniently in the acid dissolver, prior to contacting it with the catalyst. The process of this invention can be practiced by intermittent introduction of hydrogen into the bed of extended catalyst during the continuous introduction of the aqueous solution of impure terephthalic acid. The minimum of hydrogen to be introduced intermittently is, of course, an amount of hydrogen in excess of that required for reduction of the dissolved impurity so that adsorption of the excess hydrogen in the porous catalyst support can be simultaneously accomplished. Very little hydrogen is consumed by the purification process of this invention.

It is particularly advantageous according to a preferred embodiment of this invention, to impose on the aqueous solution being treated a pressure above the pressure required to maintain a liquid phase of the aqueous solution of impure terephthalic acid and dissolved hydrogen. This additional pressure prevents premature crystallization of the acid due to minor process pressure variations causing vaporization of some of the solvent. This is readily accomplished by use of an inert, non-condensable gas such as nitrogen. By "inert" gas is meant that gas which is not reactive with the terephthalic acid or the hydrogen or solvent. Nitrogen is a convenient inert gas. An additional benefit accruing from the use of nitrogen is that the dilution of the hydrogen introduced into the process of this invention provides low partial pressure of hydrogen to minimize over-hydrogenation such as, for example, saturation of aromatic nuclei.

In summary of the foregoing with respect to the hydrogenation section of the process of this invention, the hydrogen partial pressure employed in the preferred trickle system can vary in the range of from one atmosphere (14.7 psia) to 150 psi or more,

preferably in the range of 50 to 100 psi. It will be appreciated that the hydrogen driving force can be so great as to produce severe hydrogenation causing, for example, nuclear hydrogenation of the benzene ring of terephthalic acid, thereby converting some of the terephthalic acid to hexahydroterephthalic acid (cyclohexane 1,4-dicarboxylic acid), an undesirable impurity. Such severe hydrogenations should be avoided to prevent the inclusion of additional impurities to be removed in the crystallization section. Mild hydrogenation conditions should therefore be used. By the term mild hydrogenation conditions as used herein and in the appended claims is meant hydrogenation conditions which effect hydrogenation of aldehyde functional groups but which do not effect significant hydrogenation of either carboxyl groups or aromatic nuclei.

Hydrogen treating time, or space velocity, will depend on the initial terephthalic acid purity, that is, the amount of impurity to be reduced, on the desired fiber-grade specifications imposed on the purified terephthalic acid, and on other conditions of the hydrogenation such as, for example, catalyst activity. Ordinarily a treating time, i.e. contact time with the catalyst, within the range of 0.001 to 10 hours, advantageously 0.01 to 2 hours, will suffice for most operations. Although treating time is not a critical variable, it must be taken into consideration with regard to the aforementioned severe hydrogenation and its side effects.

Conditions for conducting the process of this invention can be readily ascertained, with respect to adequate vis-a-vis too severe hydrogenation, by simple tests under the conditions selected to be employed with the particular method of conducting the process of this invention. For example, by conducting these simple bench-scale tests there can be readily determined contact time in a flow system or hydrogen driving force for a percolation-type flow system, which enhance the formation of p-toluic acid and avoid formation of cyclohexane-1,4-dicarboxylic acid to a point beyond which this side product is retained in the solvent under the conditions used for crystallizing purified terephthalic acid from the hydrogen treated solution.

The hydrogenation catalyst required for the inventive process of this invention so as to convert the aldehyde carbonyl group on the 4-carboxybenzaldehyde (4-CBA) at least to a methylol group, e.g. to convert the 4-CBA to p-methylol benzoic acid, and to destroy, or otherwise render innocuous, other impurities present in the feed terephthalic acid is platinum group metal, preferably platinum or palladium, and preferably supported on adsorbent, high surface area charcoal. The platinum group metals in-

clude ruthenium, rhodium, palladium, osmium, iridium and platinum and are advantageously extended on a support such as activated carbon, e.g. adsorbent charcoal.

Though carbon-supported platinum group metal catalysts are preferred, other platinum group metal-containing catalysts may be used and reference may be made to the literature on hydrogenation and catalysts for use therein, e.g. Lohse's "Catalytic Chemistry" (Chemical Publishing Company) and particularly the section on Group VIII metal catalysts, for platinum group metal catalysts suitable for use in the liquid phase hydrogenation of the process of the present invention.

The hydrogenation catalyst used in the invention process must have sufficient hydrogenation activity to convert the aldehyde carbonyl group on the 4-carboxybenzaldehyde at least to a methylol group, e.g. p-methylol benzoic acid, and to destroy or otherwise render innocuous other impurities present in the feed terephthalic acid. A platinum group metal supported on adsorbent charcoal in the amount of 0.01—1.0 weight percent, based on total catalyst, is suitable as the hydrogenation catalyst. Advantageously, metal contents in the range of 0.05—0.5 weight percent may be used, with 0.1—0.3 weight percent being the preferred metal content for use in trickle beds of catalyst. The higher metal contents tend to produce over-hydrogenation while the lesser amounts suffer some loss in hydrogenation activity as compared with catalysts of the preferred metal content.

The adsorbent charcoal support for the platinum group metal may be any such support which has sufficient mechanical strength and surface area. It has been found that palladium-charcoal catalyst having a palladium content in the preferred range of 0.1—0.3 weight percent and also having a very high surface area in the range of 1,000—3,000 square meters per gram of catalyst are particularly well suited for use in the present invention.

The hydrogen treated solution is preferably filtered to remove any suspended solids, such as catalyst support fines and extraneous materials of about 5 microns and larger in size. The filter requirements are discussed below. The purified acid is then recovered from the filtered solution. Crystallization is a convenient and the preferred method for recovering the purified acid. Either batch or continuous crystallization may be employed in the crystallization section, and preferred embodiments of continuous and batch systems are discussed below in conjunction with the Figures shown in the accompanying drawings. Crystallized acid is recovered by centrifuging during which further purification is effected by

washing the centrifuge cake. The crystals are dried in a rotary kiln to a moisture content below 1 wt. per cent, preferably 0.02—0.06 wt. per cent, to prevent caking during subsequent storage and shipping.

Regarding materials of construction, it is advantageous to use titanium and/or titanium clad steel for metallic surfaces exposed to TA-water solutions at temperatures above 350°F. Below this temperature type 304 stainless steel is satisfactory except for the dryer tubes where type 316 stainless steel is preferred. Reinforced polyester is also suitable as an alternate material for use at temperatures below 250°F.

It has been discovered, that, surprisingly, titanium may be used for corrosion protection in contact with the acid solution, and protected from hydride formation and sloughing under reducing environment conditions in contact with hydrogen at temperatures above 350°F. and pressures above 750 psi by forming a blue oxide coating on the titanium. Such temperature and pressure conditions are encountered in the process of the present invention in such places as, for example, the dissolver preheater, dissolver, reactor, filters, crystallizers and piping. The oxide coating is conveniently formed, and maintained as required, by including a small amount of oxygen, suitably as air, in the solvent-acid slurry feed to the dissolver preheater. All titanium parts in contact with the corrosive solution under corrosion and reducing conditions are thus protected.

Turning now to the drawings, Figure 1 is simplified schematic flow plan of a preferred embodiment of the invention. Figures 2 and 3 are simplified schematic flow plans illustrating preferred embodiments of alternate continuous and batch crystallization systems, respectively. It is to be understood that these embodiments are for the purpose of illustration and are not to be regarded as a limitation of the scope of the present invention.

Referring now to Figure 1, dry crude terephthalic acid (TA) feed from source 10, such as for example a storage silo, is transferred via line 11 into crude terephthalic acid weigh hopper 12. Crude terephthalic acid is fed from the weigh hopper at a constant rate into the feed slurry tank 13 by the crude terephthalic feeder 14 which suitably can be any solids transfer feeder such as, for example, a Star feeder. The crude terephthalic feeder 14 sets the nominal feed rate to the process, but since such devices are somewhat inaccurate, the average flow rate is determined by weight as measured by the terephthalic acid weigh hopper 12. Recycled demineralized water from line 16 is added to feed slurry tank 13 on flow control to provide a slurry concentration of

crude TA in water of approximately 15—30 weight per cent, preferably 20—25 wt. per cent, and in this example about 23 weight per cent total solids. Slurry hold-up in feed slurry tank 13 of about 45 minutes at normal level is sufficient to dampen out fluctuations in the TA and water feed rates to the tank. The temperature in feed slurry tank 13 is maintained at a temperature in the range of 100—300°F., preferably about 200°F. and the pressure is conveniently near atmospheric at temperatures below the boiling point of water by venting to the atmosphere. Feed slurry tank 13 is provided with an agitator to contact the solid crude TA and the recycled water in order to maintain a uniform slurry.

Slurry is withdrawn from feed slurry tank 13 via line 17 and transferred via high-pressure pump 18 via line 19 through preheater 21. Preheater 21 is conveniently a shell-and-tube heat exchanger with one tube pass. Normally the tube-side velocity of the slurry feed is not sufficient to keep the slurry in suspension, therefore it is advantageous to mount the heat exchanger vertically and employ downflow to prevent loss of effective tube surface as a result of settling of solids. Suitable preheater outlet conditions of temperature and pressure are about 530°F. and 985 psia, respectively. The preheated slurry is passed via line 22 into dissolver 23.

The reactor feed stream is passed up-flow through the dissolver 23 which provides a residence time of approximately 20 minutes. The dissolver 23 is provided with an agitator to suspend the solids and to maintain the proper environment for a high solution rate of the crude TA into the water. A clear solution of TA in water overflows from the dissolver 23 via line 24 to the hydrogenation reactor 26 by gravity. This solution, when formed from a slurry of approximately 23 weight per cent solids, contains about 30 pounds of crude TA per 100 pounds of water at 530°F. and 985 psia. The precipitation (crystallization) point for this solution is 520°F., therefore the solution at 530°F. is only about 10° hotter than the precipitation point.

Crude terephthalic acid solution from line 24 flows continuously into a distribution pan at the top vapor space of the hydrogenation reactor 26. Overflow from the pan trickles onto the palladium-on-carbon catalyst. Flow of hydrogen is advantageously co-current to flow of the TA solution through the catalyst bed, however counter-current flow can be used. The bed is supported by a screen about equivalent to 8-gauge Tyler mesh. Catalyst particles are of such size that about 95 weight per cent are retained on this screen. However, less

than 5 per cent of the catalyst is lost from the bed because of the filtering action of the bed itself. Space is provided within the reactor vessel 26 beneath the catalyst bed for vapor disengaging and liquid-level control. Effluent vapors are vented directly to the atmosphere from this vapor space via valved line 38, and liquid level control is maintained by throttling this vented stream.

Hydrogen from source 27, preferably in admixture with an inert gas such as nitrogen, is passed via line 28 into gas preheater-saturator 29. Demineralized water from source 31 is also passed via lines 32 and 33 into gas preheater-saturator 29 wherein the hydrogen-containing gas is heated to reaction temperature and is at least partially, preferably essentially completely, saturated with water vapor. Water-saturated gas from the preheater-saturator 29 is passed via lines 34 and 36 into the top vapor space of the reactor 26 and via lines 34 and 37 into the center section of the dissolver 23 so that hydrogen is dissolved in the solution and thus is readily available to effect hydrogenation upon contact with the catalyst in reactor 26.

Prehumidification of the hydrogen-containing gas with water in preheater-saturator 29 is for the purpose of preventing localized crystallization of TA from the solution with consequent plugging which would occur if dry gas were fed into the reactor depriving the solution of water. The necessary degree of saturation of the hydrogen stream with water depends upon the equilibrium solubility of solid in the liquid, the inlet liquid temperature, and the amount of hydrogen which is consumed by the reaction. If pure hydrogen is used without inert diluent, the inlet gas stream is saturated at 530°F. and the reactor is pressure controlled, then the pressure of hydrogen in the reactor will be below 100 psia for liquid feed temperatures above 530°F. The hydrogen partial pressure in the reactor will decrease to as low as 50 psia as the temperature is increased to 535°F. The hydrogen partial pressure is greater than 100 psia for liquid feed temperatures below 530°F. but there is a smaller quantity of vent gas due to steam condensation. At a reactor temperature of 525°F. and the normal hydrogen feed rate of 0.37 moles of hydrogen flowing through the reactor per 10,000 pounds of solution feed, the vent stream will decrease to about one-fourth of the saturated gas feed rate. The normal hydrogen partial pressure in the reactor is about 100 psi with the water vapor pressure contributing an additional 885 psia. If lower hydrogen partial pressures are used, it is advantageous to include an inert gas, such as nitrogen, with the hydrogen gas feed in order to provide a pressure pad to preclude vaporization of solvent, and conse-

quent precipitation and plugging due to minor process pressure variations.

Control of liquid stream temperature is critical to stable operation. Care must be taken that the stream does not contact local cold or hot spots in pipes or vessels. In particular, if the liquid stream temperature rises above about 540°F., boiling occurs with resulting solidification of TA. Experience has indicated that it is practically impossible to redissolve this material under process conditions.

Hydrogenated solution is withdrawn from reactor 26 via line 39 and passed through reactor effluent filter 41 to remove any catalyst fines resulting from attrition of the carbon catalyst base. The filter elements are capable of removing particles of about 5 microns and larger in size. The filter elements have a minimum pressure drop corresponding to about 1 gpm/ft.²/psi with a desired maximum total pressure drop of 10 psi. Filter elements are conveniently cleaned by flushing with hot water and, if necessary, by soaking in dilute caustic. This procedure will remove from the elements any precipitated TA deposited during depressuring of the filter prior to backwashing. Prior to putting a new or cleaned filter on stream, it should be filled with water to prevent plugging due to TA solidification. Although only one filter is shown in Figure 1, it is contemplated that two or more filters may be used so that at least one filter may be in service while another is being cleaned.

Since the TA solution is highly corrosive at reactor effluent temperature, careful selection of suitable corrosion resistant elements for use in the filter is required. Filter elements from sintered titanium are satisfactory from the standpoint of corrosion, but their mechanical strength is borderline. Porous carbon beds of either loose carbon particles or bonded carbon particles have been found to be particularly suitable for this filtering application.

A suitable test to determine the effectiveness of the filter is to dissolve 5 grams of product TA in dimethylformamide and pass the resulting solution through filter paper to remove any solid fines. This filter paper is compared against a standard in which 0 represents clean filter paper and 5 is approximately gray in tone. Product TA having a rating of 0 to 1, barely distinguishable from clean filter paper, is considered satisfactory.

Filtered reactor effluent is passed via line 42 and inlet valve 82 to crystallization section 43 which is discussed in detail below and illustrated as alternate continuous and batch crystallization systems in Figures 2 and 3. In the crystallization section water is removed from the hot TA solution by high rate evaporative cooling. As a result

of both the cooling and the solvent removal, TA crystallizes from solution. Evaporated water is condensed and the condensate is withdrawn from the crystallization section and passed via line 44 into recycle solvent drum 46. The slurry of TA crystals resulting from the crystallization is withdrawn from the crystallization section via line 47 and passed into centrifuge feed tank 48 which is vented to the atmosphere by line 49. The centrifuge feed tank 48 is provided with an agitator to maintain the slurry of TA crystals in suspension. The TA slurry is passed from the centrifuge feed tank 48 via valved line 51 into one or more centrifuges 52 wherein the crystals are separated from the mother liquor and the crystals washed with fresh demineralized wash water obtained from source 31 via lines 32, 53 and 54. Wash water from the centrifuge is passed via line 55 into the recycle solvent drum 46. Mother liquor is withdrawn from the centrifuge and discarded via line 56. Purified TA crystals from the centrifuge 52 are withdrawn from the centrifuge and fed via auger feeders 57 and 57a into rotary kiln crystal dryer 58 which is heated by steam from source 59. Inert gas, such as nitrogen, from source 61 may be passed via line 62 into and through drying kiln 58 to assist in removing moisture from the TA crystals. The product is dried to a moisture content of about 0.05 wt. per cent water. Inert gas and water vapor containing some TA fines are withdrawn from the kiln 58 and passed via line 63 into dryer-scrubber 64 wherein the gases are washed countercurrently by a stream of fresh demineralized water. The water is introduced to the dryer-scrubber 64 from source 31 via lines 32, 53 and 66 and descends countercurrently to the rising gas stream removing TA fines from the gases which are vented from the scrubber via line 67. Water containing TA fines is withdrawn from the dryer-scrubber 64 and passed via line 68, pump 69 and line 71 into the recycle solvent drum 46. The solvent drum is vented to the atmosphere via line 72. An agitator is provided within the solvent recycle drum 46 in order to maintain undissolved particles of TA, primarily from scrubber 64, in suspension. Recycle solvent water is withdrawn from the recycle solvent drum 46 and passed via line 73, pump 74 and line 16 into the feed slurry tank 13. Dried purified TA product is withdrawn from the drying kiln 58 and passed via conveyor 76 to storage.

The purified product contains less than 25 ppm by weight of 4-CBA and also less than 95 ppm by weight of p-toluic acid. The T.E.G. color of the product is below 50 which is well below the suitable maximum of 150 and the desirable maximum of 100.

The 4-carboxybenzaldehyde content was determined by dissolving the terephthalic acid in dilute alkali and buffering to a pH of 9, followed by polarographic analysis to determine "4-CBA" content. Color, expressed as "T.E.G. Color" (triethylene glycol color) was determined by esterifying 4.0 grams of the terephthalic acid with 28.4 cc of triethylene glycol at 500°F., followed by dilution to 50 vol. % with isopropanol and comparing the resultant solution color with American Public Health Association (APHA) standards.

Reference is now made to Figure 2 which is a simplified schematic flow plan illustrating a preferred embodiment of a continuous crystallization system for use in the process of the present invention. The first crystallizer 81 receives filtered reactor effluent at 530°F. and 985 psia via valved line 42 (Figure 1). Flashing occurs across the inlet valve 82 to 725 psia and a temperature of about 507°F. Vapor released during flashing is passed via valved line 83 and line 84 into the crystallizer condenser 86. The hold-up time in the first crystallizer 81 is about 1:1 hours. Approximately 30 per cent of the dissolved terephthalic acid is precipitated at a rate of about 0.15 lbs. TA/min./100 lbs. H₂O. The first crystallizer 81 is provided with an agitator to suspend the terephthalic acid crystals in the form of a slurry. Slurry is withdrawn from the first crystallizer 81 and passed via valved line 87 into the second crystallizer 88. Slurry concentration in the effluent from the first crystallizer is approximately 7.1 wt. per cent solids.

The second crystallizer 88 is operated at 115 psia and a temperature of about 338°F. Vapor resulting from the further flashing of solvent in the second crystallizer 88 is passed via valved line 89 and line 84 into the crystallizer condenser 86. Hold-up time in the second crystallizer 88 is about 1.2 hours. The terephthalic acid crystallization rate in the second crystallizer is about 0.4 lbs. TA/min./100 lbs. H₂O. Effluent slurry concentration from the second crystallizer is approximately 29.1 wt. per cent solids. The second crystallizer 88 is also provided with an agitator to maintain the TA crystals in slurry suspension.

Slurry is withdrawn from the second crystallizer 88 and passed via valved line 91 into the third crystallizer 92 which operates at a temperature of about 228°F. and 20 psia with a liquid hold-up time of about 1.1 hours. Nearly all of the terephthalic acid, about 96.6 wt. per cent, is crystallized under the operating conditions in the third crystallizer. Flashed solvent vapors are withdrawn from the third crystallizer 92 and passed via valved line 93 and line 84 into the crystallizer condenser 86. The third

crystallizer 92 is also provided with an agitator to maintain the crystallized TA in suspension in the form of a slurry. Condensate from crystallizer condenser 86 is passed via line 44 into recycle solvent drum 46 (Figure 1). Slurry containing about 32.5 wt. per cent solids is withdrawn from the third crystallizer 92 and passed via valved line 47 into centrifuge feed tank 48 (Figure 1).

In reference to Figure 3, which is a simplified schematic flow plan illustrating a preferred embodiment of a batch crystallization system, it is pointed out that this batch system and the previously discussed continuous crystallization system are alternative crystallization sections for use in the process of this invention as illustrated in Figure 1. Referring now to Figure 3, effluent hold tank 101 receives filtered reactor effluent continuously via valved line 42 (Figure 1) under temperature and pressure conditions of 530°F. and 985 psia, respectively. A small amount of solvent flashing occurs across the inlet valve 82 to about 885 psia with only a minor decrease in stream temperature. Essentially all of the dissolved hydrogen is released from the liquid in this vessel. Solvent vapor and gas is passed via valved line 102 and line 84 into crystallizer condenser 86.

While three crystallizer vessels are illustrated in Figure 3, it should be understood that any number, one or more, of crystallizer vessels may be employed. Hot TA solution is alternatively passed from effluent hold tank 101 into each of the crystallizers A, B and C via line 103 and valved manifold line 104. High pressure steam, nominally 1,000 psi, is used to pressure each crystallizer prior to its fill cycle in order to prevent flashing of the hot solution into a nearly empty crystallizer vessel causing plugging of the inlet line. This steam is introduced from source 106 via line 107 and valved manifold 108 into the vapor space of each crystallizer prior to its fill cycle. High pressure steam is also fed via valved line 109 into effluent hold tank 101 during its dumping cycle to displace charged solution therefrom preventing excessive solvent vaporization. This steam is then vented via valved line 102 and line 84 into the crystallizer condenser 86 during the hold tank's filling cycle.

The crystallization cycle of each crystallizer includes the following: vessel heating and pressurization with steam, filling, cooling and crystallizing, and dumping. The cooling period is divided into three regions of limiting conditions:

- 1.) Crystallization rate,
- 2.) Vaporization rate, and
- 3.) Vapor loading and defoaming rate.

The rate of cooling during the first 15 minutes of the cooling and crystallizing period is limited to a crystallization rate of 1.5 lbs. TA/100 lbs. H₂O/min. by controlling the rate of water vapor flow from the crystallizer. The final cooling period is self limited by the rate of release of water vapor from the liquid surface at the lower temperatures existing at the time. The maximum crystallization rate has been set at 1.5 lbs. TA/100 lbs. H₂O/min. in order to prevent shock cooling with the consequent formation of an excessive number of infinitesimal crystals and, more importantly, to prevent inclusion within the TA crystals of paratoluic acid and other impurities, which inclusion has been found to occur excessively at crystallization rates such as result when the solvent (water) vapor is flashed without controlling the evaporation rate such as can be conveniently done by throttling the vapor outlet line.

Crystallizer agitators are provided in each of the crystallizer vessels to maintain slurry suspension during the cooling and dumping cycles. Vapors from each of the crystallizers are passed via manifold line 110 and line 84 into crystallizer condenser 86.

Crystallizer effluent is transferred via manifold line 111 into the crystallizer hold tank 112 by pressuring the crystallizers with 150 psig steam from a source not shown. This transfer is performed under steam pressure to avoid flashing in manifold line 111 with consequent plugging. The effluent is flashed across the inlet valve 113 to atmospheric pressure in crystallizer hold tank 112, which pressure is maintained by venting to the atmosphere via valved line 114. Crystallizer hold tank 112 is provided with an agitator to maintain the TA crystals in slurry suspension. Slurry is pumped from the crystallizer hold tank 112 via line 116, transfer pump 117 and valved line 47 into centrifuge feed tank 48 (Figure 1).

Vapor loading and defoaming rate are handled according to conventional design techniques.

In Patent No. 994,769 there is described and claimed a process for producing fibre-grade terephthalic acid from an impure terephthalic acid containing 4-carboxybenzaldehyde and other impurities, which process comprises catalytically hydrogenating said impure terephthalic acid in an aqueous solution at elevated temperature and pressure of a time sufficient to effect substantial reduction of said 4-carboxybenzaldehyde, and recovering a terephthalic acid of fibre-grade quality containing substantially less of said 4-carboxybenzaldehyde and other impurities.

In our copending Patent Application No. 21655/66 (Serial No. 1,152,576) there is described and claimed a process for purifying

an aromatic polycarboxylic acid containing undesirable aldehyde and other impurities, which process comprises contacting a solution of the impure acid in water in the liquid phase at an elevated temperature and pressure and in the presence of prehumidified hydrogen-containing gas with a platinum group metal-containing catalyst under mild hydrogenation conditions (as defined therein) and recovering purified acid containing substantially less of the said undesirable impurities.

Further in our copending Patent Application No. 21656/66 (Serial No. 1,152,577) there is described and claimed a process for purifying an aromatic polycarboxylic acid containing undesirable aldehyde and other impurities, which process comprises contacting a solution of the impure acid in the liquid phase at an elevated temperature and pressure in demineralized water and in the presence of hydrogen with a palladium-containing catalyst under mild hydrogenation conditions (as defined therein) and recovering purified acid containing substantially less of said undesirable impurities.

WHAT WE CLAIM IS:—

1. A process for purifying aromatic polycarboxylic acid containing undesirable aldehyde and other impurities, comprising contacting a solution of the impure acid in water in the liquid phase at an elevated temperature and pressure and in the presence of hydrogen with a platinum group metal-containing catalyst under mild hydrogenation conditions (as hereinbefore defined) and recovering from the resulting solution purified acid containing substantially less of said undesirable impurities by crystallization employing controlled-rate evaporative cooling.

2. A process according to claim 1, wherein said polycarboxylic acid is an acid produced by liquid phase catalytic oxidation of a polyalkyl aromatic hydrocarbon.

3. A process according to claim 2, wherein said polycarboxylic acid is a dicarboxylic acid produced by liquid phase catalytic oxidation of a dialkyl aromatic hydrocarbon.

4. A process according to claim 3, wherein said acid is terephthalic acid, said dialkyl aromatic hydrocarbon is paraxylene, and said aldehyde is 4-carboxybenzaldehyde.

5. A process according to any preceding claim, wherein the crystallization is carried out batchwise.

6. A process according to any one of claims 1 to 4, wherein the crystallization is carried out continuously.

7. A process according to claim 5, wherein said crystallization is effected in a batch crystallization zone and said zone is first pressured with steam to at least the pressure corresponding to the vapor pres-

sure of water at the temperature of the solution feed to said zone prior to introducing said feed to said zone.

8. A process according to claim 7, wherein said crystallization is effected at a water solvent evaporation rate producing crystallization of acid at a rate below 1.5 lbs of acid/100 lbs of water/minute.

9. A process according to claim 6, wherein said crystallization is effected in a continuous crystallization system employing a plurality of crystallization zones in series flow arrangement, each of said zones being maintained at lower temperature and pressure conditions than the preceding one.

10. A process according to claim 1 or claim 1 and any one of claims 5 to 9, wherein the impure acid is crude aromatic polycarboxylic acid which has been obtained by catalytic liquid phase oxidation of a polyalkyl aromatic hydrocarbon with molecular oxygen in the presence of a heavy metal oxidation catalyst, said aromatic hydrocarbon having at least two nuclear alkyl hydrocarbon substituents whose carbon attached to the nuclear aromatic carbon has at least one hydrogen atom, and which has an aromatic polycarboxylic acid content of at least 99.0% by weight and has as its principal impurity a carboxy aromatic aldehyde corresponding to said aromatic polycarboxylic acid.

11. A process according to claim 10, wherein said contacting takes place at a temperature of from 450 to 600°F for a time sufficient to effect substantial reduction of said aldehyde, and purified acid containing less than 125 ppm of said aldehyde is recovered.

12. A process according to claim 10 or 11, wherein said crude aromatic polycarboxylic acid in a crude aromatic dicarboxylic acid obtained by catalytic liquid phase oxidation of a dialkyl aromatic hydrocarbon with molecular oxygen in the presence of a heavy metal oxidation catalyst and wherein said principal impurity is a monocarboxy aromatic aldehyde corresponding to said aromatic dicarboxylic acid.

13. A process according to claim 12, wherein said crude aromatic dicarboxylic acid is crude terephthalic acid, said dialkyl hydrocarbon is paraxylene and said monocarboxy aromatic aldehyde is 4-carboxybenzaldehyde.

14. A process according to any one of the preceding claims, wherein said platinum group metal is platinum.

15. A process according to any one of claims 1 to 13, wherein said platinum group metal is palladium.

16. A process for the purification of an aromatic polycarboxylic acid according to claim 1 and substantially as herein described.

17. A purified aromatic polycarboxylic acid when produced by a process according to any one of the preceding claims.

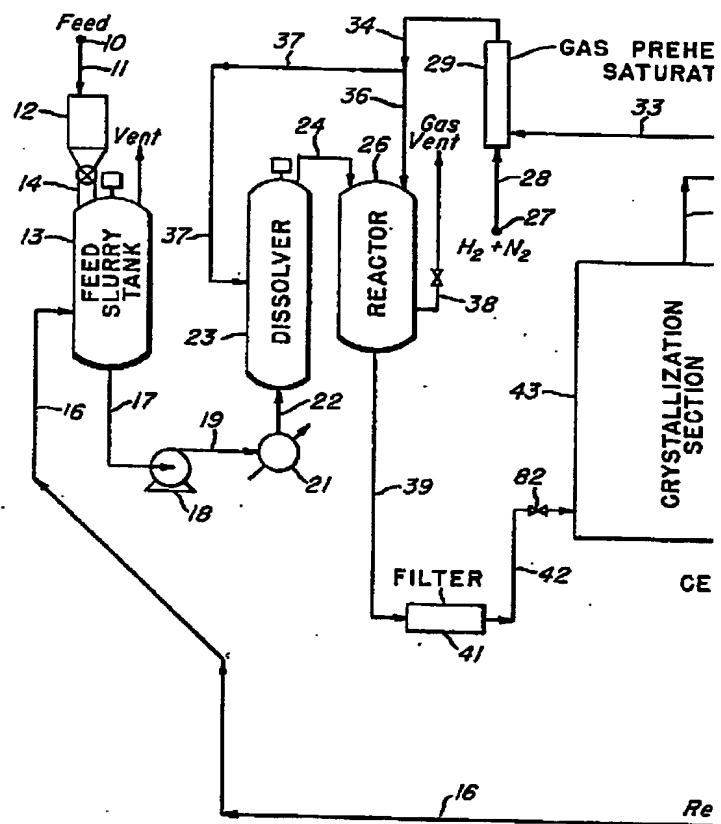
5 18. Fiber-grade terephthalic acid when produced by a process according to any one of claims 1 to 16.

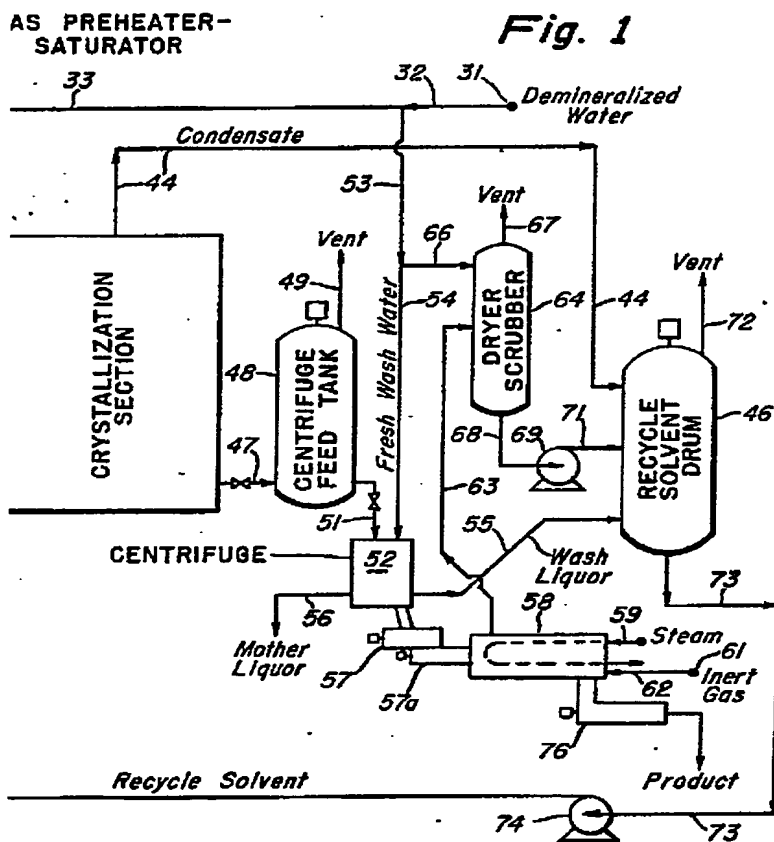
19. Apparatus substantially as herein described with reference to the accompanying drawings when used for the purification

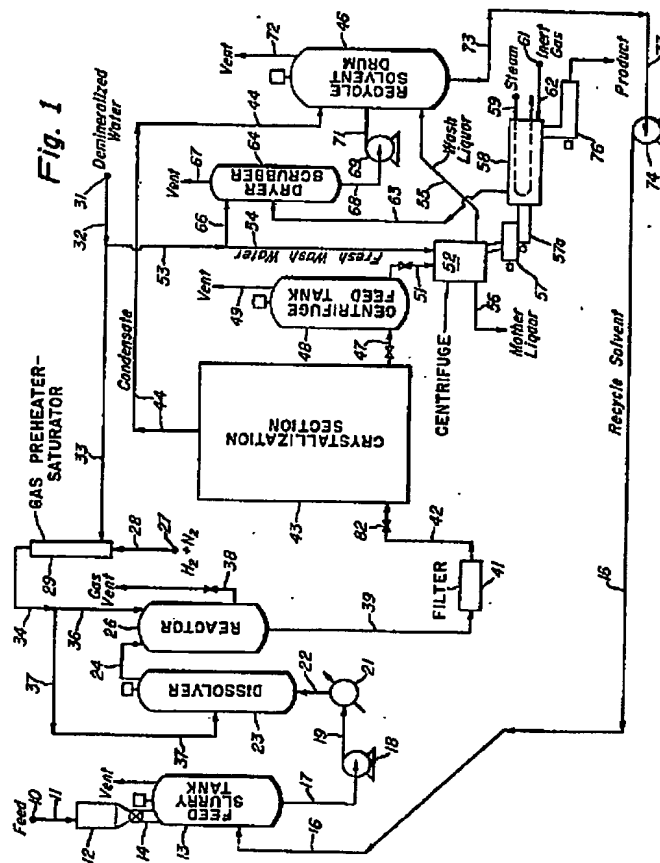
of an aromatic polycarboxylic acid by the process claimed in any one of claims 1 to 16. 10

MATHYS & SQUIRE,
Chartered Patent Agents,
10 Fleet Street,
London, E.C.4.
Agents for the Applicants.

Printed for Her Majesty's Stationery Office by Burgess & Son (Abingdon), Ltd.—1969.
Published at The Patent Office, 25 Southampton Buildings, London, W.C.2,
from which copies may be obtained.







CONTINUOUS CRYSTALLIZATION SECTION

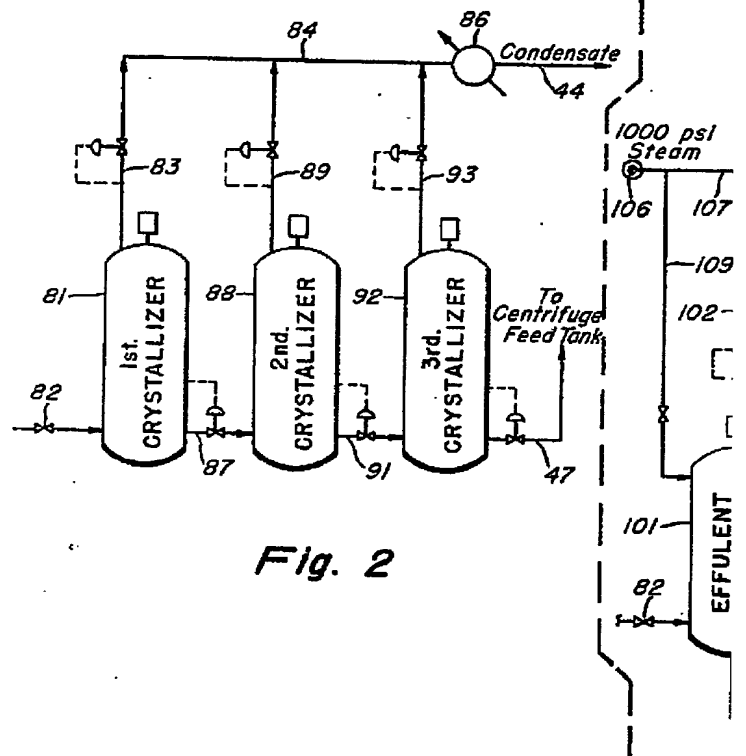


Fig. 2

1152575

COMPLETE SPECIFICATION

2 SHEETS

This drawing is a reproduction of
the Original on a reduced scale

Sheet 2

BATCH CRYSTALLIZATION SECTION

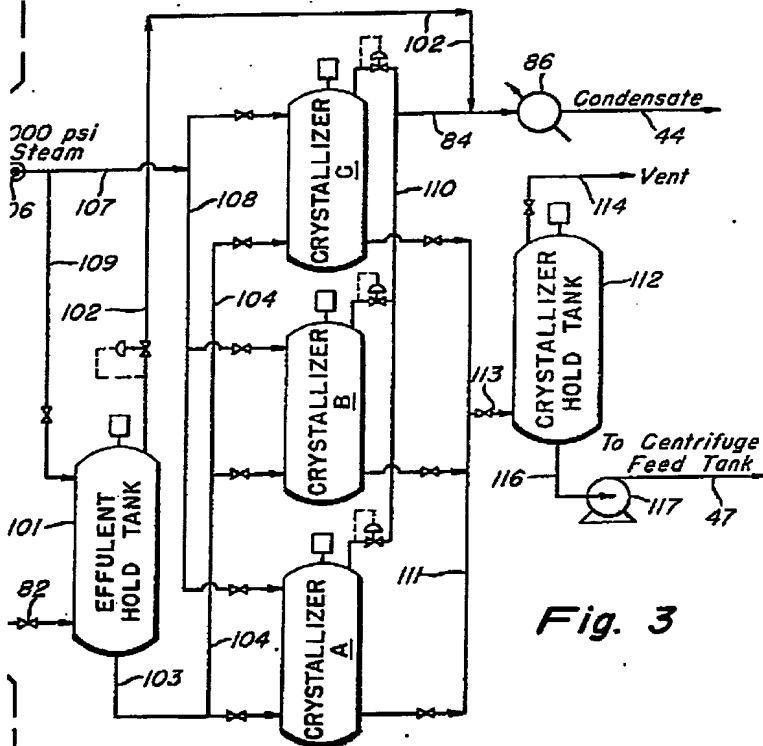


Fig. 3

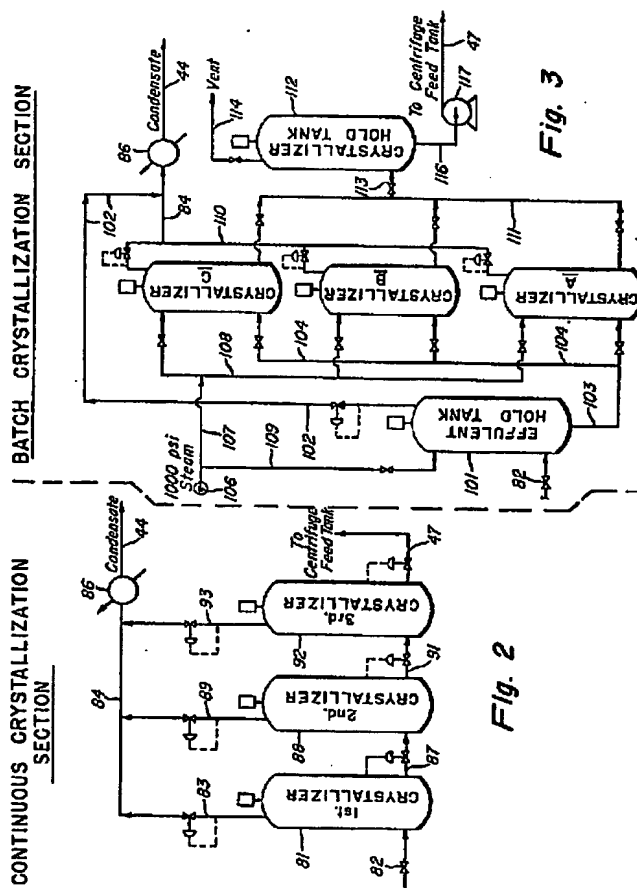


Fig. 2

Fig. 3